

action of acrolein on urea, was that containing the residue from one molecule of urea, the compound $\text{CO}(\text{NH})_2 \cdot \text{C}_3\text{H}_4$.

XIII. NOTE ON AN EARTHY FERRIC SULPHATE FROM ARKANSAS.

BY CHAS. E. WAIT.

Several weeks ago a sample of earth was sent to me for examination, from Southern Arkansas. It is a friable, yellowish substance, partly soluble in water, streak uncolored, taste astringent.

Upon analysis I find it contains the following :

	PER CENT.
Insoluble in cold water.....	12.61
Soluble in cold water.....	87.39
	<hr style="width: 100%;"/>
	100.00
The soluble portion consists of	
$\text{Fe}_2 \text{O}_3$	21.82
SO_3	33.81
H_2O	31.76
	<hr style="width: 100%;"/>
	87.39
By assuming the following arrangement,	
$\text{Fe}_2 \text{O}_3$	21.82
SO_3	33.81
H_2O combined.....	22.41
$\text{H}_2 \text{O}$ hygroscopic.....	9.35
	<hr style="width: 100%;"/>
	87.39

we are able to deduce a formula consistent with that of the native persulphate of iron, viz.: $\text{Fe}_2 (\text{SO}_4)_3 (\text{H}_2\text{O})_9$.

If all the water be considered as belonging to the compound, the formula will be $\text{Fe}_2 (\text{SO}_4)_3 (\text{H}_2\text{O})_{13}$.

It is said an unlimited supply of this substance may be obtained.